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SI unit representing

$6.02 \times 10^{23}$

representative

particles of a

substance. the

temperature and

pressure at which one

mole of gas occupies a

volume of 22.4 L. equal

volumes of gases at

the same temperature

and pressure contain

equal numbers of

particles.

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Chapter 10 "Chemical Quantities". the SI unit representing  $6.02 \times 10^{23}$  representative particles of a substance. the temperature and pressure at which one mole of gas occupies a volume of 22.4 L. equal volumes of gases at the same temperature and pressure contain equal numbers of particles.

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10.2 Mole to Mole &  
Mole to Volume  
Relationships \*For all  
the problems on this  
page, first find the  
molar mass of the  
compound.

1. What is  
the mass of 9.45 mol of  
aluminum oxide? mol A  
ILO + 3 (I A GO q  
63,5ù?- 2. What is the  
mass of  $4.52 \times 10^{-3}$  mol  
of ethylbenzene,

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$C_6H_5CH_2CH_3$ ?

\*Ethylbenzene is a hydrocarbon that is produced by burning coal.

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CHAPTER 10: Chemical  
Quantities BASICS: •

The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that

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substance • The mole was founded by a scientist named Avagadro, and he decided to use the

## **CHAPTER 10: Chemical Quantities**

10. Look at Figure 10.16 and Table 10.3. Name three compounds that have an empirical formula of CH. 11. Fill in the labels on the diagram below.

MICROSCOPIC  
INTERPRETATION

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MACROSCOPIC  
INTERPRETATION SO<sub>3</sub>  
molecule 1 mol SO<sub>3</sub>  
SO<sub>3</sub> composed of  
composed of S atom  
and sulfur atoms ( 1023) oxygen atoms 3  
atoms and CHAPTER  
10, Chemical  
Quantities(continued)

## **SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)**

Chemical Quantities

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Quantities 243 Section  
Review Objectives •

Convert the mass of a  
substance to the  
number of moles of a  
substance, and the  
number of moles of a  
substance to mass •

Calculate the volume  
of a quantity of gas at  
STP Vocabulary •

Avogadro's hypothesis  
• standard

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temperature and  
pressure (STP) • molar  
volume Key Equations  
• mass (grams)  
number of moles

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Chapter 10 Chemical  
Quantities 243 Section  
Review Objectives •  
Convert the mass of a  
substance to the  
number of moles of a  
substance, and the  
number of moles of a



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substance to mass •

Calculate the volume  
of a quantity of gas at  
STP Vocabulary •

Avogadro's hypothesis  
• standard

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pressure (STP) • molar  
volume Key Equations

• mass (grams)  
number of moles

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chapter 2/chapter

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10/chapter 11 chemical quantities. day: class activity: assignment: worksheets/links/podcasts: wed. 11/20. go over test

## **CHAPTER 2/CHAPTER 10/CHAPTER 11**

Chapter 10 "Chemical Quantities" Use this activity to check your knowledge of the vocabulary presented in this chapter

**Quia - Chapter 10**

*Page 18/21*

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10, that can be made from  $1.09 \times 10^4$  kilograms of phosphorus in the second of these three reactions. The answer is  $2.50 \times 10^4$  kg P 40 10. We used the following steps:  
Balance chemical equations. (Section 4.1.) Write or identify the definitions of solution, solute, and solvent. (Section 4.2)

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Given a description of  
a solution, identify the

## **Chapter 10 Chemical Calculations and equations**

Chemical Quantities

THE MOLE AND

QUANTIFYING MATTER

10.1 The Mole: A

Measurement of Matter

Essential

Understanding The

mole represents a  
large number of very  
small particles.

Reading Strategy

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